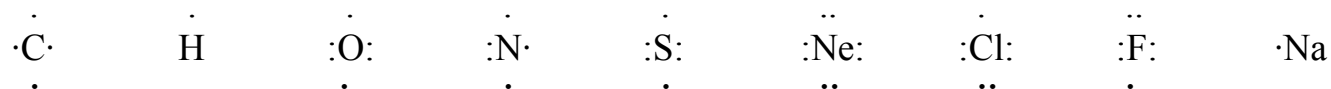


Lewis Dot Structure Answer sheet

Write the Lewis dot structures for each of the following



Complete the following table. Given are italicized.

FORMULA	NAME	LEWIS DOT STRUCTURE
<i>H<sub>2</sub></i>	<i>hydrogen gas (dihydrogen)</i>	<i>H—H</i>
<i>F<sub>2</sub></i>	<i>flourine gas (diflourine)</i>	$\overset{\cdot\cdot}{\text{:F}}\text{—}\overset{\cdot\cdot}{\text{F}}\text{:}$ $\overset{\cdot\cdot}{\cdot\cdot}$
<i>Cl<sub>2</sub></i>	<i>chlorine gas (dichlorine)</i>	$\overset{\cdot\cdot}{\text{:Cl}}\text{—}\overset{\cdot\cdot}{\text{Cl}}\text{:}$ $\overset{\cdot\cdot}{\cdot\cdot}$
<i>N<sub>2</sub></i>	<i>nitrogen gas (dinitrogen)</i>	$\text{:N}\equiv\text{N:}$
<i>O<sub>2</sub></i>	<i>oxygen gas (dioxygen)</i>	$\overset{\cdot\cdot}{\text{:O}}=\overset{\cdot\cdot}{\text{O}}\text{:}$
<i>NH<sub>3</sub></i>	<i>ammonia</i>	$\overset{\cdot\cdot}{\text{H}}\text{—}\overset{\cdot\cdot}{\text{N}}\text{—}\overset{\cdot\cdot}{\text{H}}$   $\overset{\cdot\cdot}{\text{H}}$
<i>H<sub>2</sub>O</i>	<i>dyhydrogen oxide (water)</i>	$\overset{\cdot\cdot}{\text{H}}\text{—}\overset{\cdot\cdot}{\text{O}}\text{—}\overset{\cdot\cdot}{\text{H}}$ $\overset{\cdot\cdot}{\cdot\cdot}$
<i>CO<sub>2</sub></i>	<i>carbon dioxide</i>	$\overset{\cdot\cdot}{\text{:O}}=\overset{\cdot\cdot}{\text{C}}=\overset{\cdot\cdot}{\text{O}}\text{:}$
<i>CH<sub>4</sub></i>	<i>methane</i>	$\overset{\cdot\cdot}{\text{H}}$   $\overset{\cdot\cdot}{\text{H}}\text{—}\overset{\cdot\cdot}{\text{C}}\text{—}\overset{\cdot\cdot}{\text{H}}$   $\overset{\cdot\cdot}{\text{H}}$
<i>C<sub>2</sub>H<sub>2</sub></i>	<i>acetylene</i>	$\text{H—C}\equiv\text{C—H}$
<i>C<sub>2</sub>H<sub>4</sub></i>	<i>ethylene (ethene)</i>	$\text{H—C}=\text{C—H}$        $\overset{\cdot\cdot}{\text{H}} \quad \overset{\cdot\cdot}{\text{H}}$
<i>CS<sub>2</sub></i>	<i>carbon disulfide</i>	$\overset{\cdot\cdot}{\text{:S}}=\overset{\cdot\cdot}{\text{C}}=\overset{\cdot\cdot}{\text{S}}\text{:}$
<i>CCl<sub>4</sub></i>	<i>carbon tetrachloride</i>	$\overset{\cdot\cdot}{\text{:Cl}}\text{:}$   $\overset{\cdot\cdot}{\text{:Cl}}\text{—}\overset{\cdot\cdot}{\text{C}}\text{—}\overset{\cdot\cdot}{\text{Cl}}\text{:}$ $\overset{\cdot\cdot}{\cdot\cdot}$   $\overset{\cdot\cdot}{\cdot\cdot}$   $\overset{\cdot\cdot}{\text{:Cl}}\text{:}$ $\overset{\cdot\cdot}{\cdot\cdot}$
<i>HF</i>	<i>hydrogen flouride</i>	$\overset{\cdot\cdot}{\text{H}}\text{—}\overset{\cdot\cdot}{\text{F}}\text{:}$ $\overset{\cdot\cdot}{\cdot\cdot}$

$H_2S$	dihydrogen sulfide (hydrogen sulfide)	$\begin{array}{c} \cdot\cdot \\ \text{H}-\text{S}-\text{H} \\ \cdot\cdot \end{array}$
$NI_3$	nitrogen triiodide	$\begin{array}{c} \cdot\cdot \quad \cdot\cdot \quad \cdot\cdot \\ :\text{I}-\text{N}-\text{I}: \\ \cdot\cdot \quad   \quad \cdot\cdot \\ \quad \cdot\cdot \\ \quad \cdot\cdot \end{array}$
$H_2Se$	<i>dihydrogen selenide</i>	$\begin{array}{c} \cdot\cdot \\ \text{H}-\text{Se}-\text{H} \\ \cdot\cdot \end{array}$